

NAME: _____

Reminders:

1. In a neutral atom the number of protons equals the number of electrons.
2. An atom can NEVER gain or lose protons
3. The number of protons equals the atomic number

Ion Practice Set

1. What is an ion?
2. What does the number next to the ions signify?

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	NUMBER OF ELECTRONS LOST OR GAINED
ex	Fluorine	F ⁻	9	10	gained one
1			53	54	
2			16		gained two
3	potassium				lost one
4		Ca ⁺²			
5			35	36	
6		Sr ⁺²			
7		H ⁺			
8			8		gained two
9			12		lost two
10	aluminum			10	
11			34	36	
12		H ⁻			
13	lithium				lost one
14		Rb ⁺			
15			17	18	

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Isotopes Practice Set

1. What is an isotope?
2. What does the number next to isotopes signify?
3. How can you tell isotopes apart?

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chromium-58	Chromium-63
# of protons		
# of neutrons		
# of electrons		

	Carbon-12	Carbon-16
# of protons		
# of neutrons		
# of electrons		

	Nitrogen-15	Nitrogen-20
# of protons		
# of neutrons		
# of electrons		

	Sulfur-23	Sulfur-25
# of protons		
# of neutrons		
# of electrons		

Fill in the isotope names and any missing information, including isotope numbers from the chart. Use your periodic table and the information provided.

	Iodine-	Iodine-
# of protons		
# of neutrons	32	35
# of electrons		

	Iron-	Iron-
# of protons		
# of neutrons	27	30
# of electrons		

# of protons		
# of neutrons	113	111
# of electrons	55	

# of protons	32	
# of neutrons	30	32
# of electrons		