

### Ionic Bonds 101 Video Review

1. Draw the Lewis dot structures for Beryllium and Fluorine below.
2. Negative ion is called an \_\_\_\_\_, has a \_\_\_\_\_ charge and has \_\_\_\_\_ an electron.
3. Positive ion is called a \_\_\_\_\_, has a \_\_\_\_\_ charge and has \_\_\_\_\_ an electron.
4. What two things do elements in the same column have in common? \_\_\_\_\_  
What does this help determine? \_\_\_\_\_
5. What does it mean if an element has a -2 charge? \_\_\_\_\_  
What does it mean if an element has a +1 charge? \_\_\_\_\_
6. What is an ionic bond? \_\_\_\_\_  
\_\_\_\_\_  
Ionic bonds are formed between a \_\_\_\_\_ and a \_\_\_\_\_.
7. Draw the Lewis dot structure ionic bond for Sodium (Na) and Chlorine (Cl) below.
  
8. How does sodium get a +1 charge? \_\_\_\_\_  
How does chlorine get a -1 charge? \_\_\_\_\_
9. What causes the bond between Na and Cl? \_\_\_\_\_
10. Draw the Lewis dot structure ionic bond for  $\text{AlCl}_3$  (Aluminum Chloride) below.
  
11. What makes  $\text{AlCl}_3$  an ionic bond? \_\_\_\_\_
12. What does Al do in the chemical bond? \_\_\_\_\_  
What does Cl do in the chemical bond? \_\_\_\_\_
13. What is the cross method? \_\_\_\_\_
14. Draw the Lewis dot structure ionic bond for BeO (Beryllium Oxide) below.

Practice making the following ionic bonds

15. Potassium + Fluorine      16. Sodium + Sodium + Oxygen      17. Chlorine + Calcium + Chlorine